

Soluciones Ana Sevillega

Calcular la masa molecular d:

$$\text{Fe}_2\text{O}_3 = 56 \cdot 2 + 16 \cdot 3 = 143 \text{ umas.}$$

$$\text{Cl}_2\text{O}_3 = 35 \cdot 2 + 16 \cdot 3 = 118 \text{ umas.}$$

$$\text{SO}_3 = 32 + 16 \cdot 3 = 80 \text{ umas.}$$

$$\text{AlH}_3 = 27 + 1 \cdot 3 = 30 \text{ umas}$$

$$\text{PbH}_4 = 207 + 1 \cdot 4 = 211 \text{ umas.}$$

$$\text{Ba}(\text{OH})_2 = 137 + (16 + 1) \cdot 2 = 171 \text{ umas.}$$

$$\text{NaClO}_4 = 27 + 35 + 16 \cdot 4 = 126 \text{ umas.}$$

$$\text{K}_2\text{SO}_4 = 40 \cdot 2 + 32 + 16 \cdot 4 = 176 \text{ umas.}$$

SESION 3 TEMA 7

datos

$$\text{HFe} = 56 \mu$$

$$\text{HO} = 16 \mu.$$

$$\text{HCl} = 35 \mu.$$

$$\text{Hg} = 207 \mu$$

$$\text{HAl} = 27 \mu.$$

$$\text{HH} = 1 \mu.$$

$$\text{HCl} = 35 \mu.$$

$$\text{HBr} = 137 \mu.$$

$$\text{HNa} = 23 \mu.$$

$$\text{HK} = 40 \mu.$$